

SYNTHESIS AND CHARACTERIZATION OF GADOLINIUM-BASED NANOCRYSTALS FOR THEIR POTENTIAL APPLICATION AS TRIMODAL IMAGING CONTRAST AGENTS

HAROON UR RASHID ^a, MUHAMMAD YASEEN ^b, ALI RIAZ ^{a*}, KAMIN KHAN ^a, HANIF UR RAHMAN ^a,
KHALID KHAN ^c, NASIR AHMAD ^c AND MARCO ANTONIO UTRERA MARTINES ^{d*}

^aDepartment of Chemistry, Sarhad University of Science & Information Technology, Peshawar, Khyber Pakhtunkhwa, Pakistan.

^bInstitute of Chemical Sciences, University of Peshawar, Peshawar, Khyber Pakhtunkhwa, Pakistan.

^cDepartment of Chemistry, Islamia College University, Peshawar, Khyber Pakhtunkhwa, Pakistan.

^dInstitute of Chemistry, Federal University of Mato Grosso do Sul, Campo Grande, MS, Brazil.

ABSTRACT

Lanthanide-doped gadolinium nanomaterials have attracted considerable attention due to their promising applications in biological imaging. Sodium gadolinium fluoride (β -NaGdF₄) nanomaterials act as perfect host for doping of luminescent lanthanide ions. Due to heavy nature of gadolinium, it can strongly absorb X-ray radiations. Therefore, Gd-based nanomaterials are applied as contrast agents for X-ray tomography (CT). Presence of seven unpaired electrons in Gd³⁺ ion, its large magnetic moment and long electronic relaxation time makes it an ideal candidate to enhance water proton relaxation rates. These characteristics make Gd-based nanomaterial useful for their potential application as MRI contrast agents. In this work, sodium gadolinium fluoride is doped with ytterbium and erbium to produce β -NaGdF₄:Yb/Er as core nanocrystals. They are subsequently coated with sodium gadolinium fluoride doped with neodymium (β -NaGdF₄:Nd) to produce β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd core-shell nanocrystals. Powder X-ray diffraction (XRD) results indicated that β -NaGdF₄ (host), β -NaGdF₄:Yb/Er (core) and β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd (core-shell) nanocrystals all exist in hexagonal crystalline phase. While Dynamic Light Scattering (DLS) results indicated that the particle size of the synthesized nanocrystals was in range of 12 to 27 nm. FTIR results confirmed that the synthesized nanocrystals were stabilized by oleic acid. Nanocrystals reported in this work are expected to be useful trimodal contrast agents for photoluminescence, X-ray tomography and Magnetic Resonance Imaging (MRI).

Keywords: lanthanide, nanocrystals, trimodal, gadolinium, contrast agents.

INTRODUCTION

Recently, rapid research developments in the field of medicine and life science have also driven the biomedical research. This emerging and growing research field comprises of a number of bioassay approaches such as X-ray Computed Tomography (also called Computed Tomography, CT), Single-Photon Emission Computed Tomography (SPECT) [1-4], Positron Emission Tomography (PET) [5, 6] and Magnetic Resonance Imaging (MRI) [7-9]. These tomographic imaging methods have high penetration power. Therefore, they play a key role in biomedical imaging. However, these techniques have some unavoidable drawbacks such as limited spatial resolution, long scan acquisition time and application of strong ionizing radiations (except for MRI) [10-13]. Furthermore, these modalities differ from each other in spatial and temporal resolution, imaging depths, anatomical and molecular details and the necessary material properties of contrast agents essential to improve the quality of images in terms of their visibility. In many cases, it is beneficial to use several complimentary imaging techniques for rapid and precise diagnosis. Since majority of these imaging modalities apply exogenous contrast agents to improve the signal-to-noise ratio, therefore, the development of multimodal contrast agents is highly desirable to obtain images of better quality. Such agents have proved more useful for patient care.

Moreover, they also reduce the cost and increase the safety by restricting the number of contrast agents administration needed for imaging purposes [14]. The idea of designing multimodal contrast agents provides the prospective to integrate the benefits of different modalities while at the same time overcoming the shortcomings of one another [15, 16]. In addition, multimodal contrast agents can be very useful for the technical development of emerging imaging techniques [17-19]. Current research indicate that nanomaterial-based multimodal imaging probes have been developed for accurate imaging and diagnosis while applying different imaging modalities. Recently, lanthanide metal ions doped with upconverting (Excitation at lower energy than emission) nanomaterial have been investigated for their potential applications in biomedical imaging. Such nanomaterials possess high photostability and show neither photobleaching nor photoblinking [20]. Additionally, upconverting nanomaterials are excited by near-infrared (NIR) radiations which ensure optical imaging of deeper tissues

with significantly controlled autofluorescence, and create enhanced image contrast [21-23]. β -NaGdF₄ host nanomaterial doped with lanthanide ions combination of Yb³⁺, Er³⁺ or Yb³⁺, Tm³⁺ have exhibited significant upconversion luminescence [24]. Gadolinium is a heavy metal of high atomic number which strongly absorbs X-rays. Its trivalent ion (Gd³⁺) causes enhancement of water proton relaxation rate due to maximum number (seven) of unpaired electrons (among all lanthanide ions), a long electronic relaxation time ($\sim 10^{-9}$ s) and a large magnetic moment (7.9 BM). These properties make Gadolinium the best candidate for applications as luminescence, CT and MRI contrast agents [11, 25, 26]. Upconverting nanomaterial based multimodal contrast agents can serve for fluorescence imaging, CT and MRI modalities. Such contrast agents possess the advantages of high sensitivity (fluorescence imaging) as well as the potential of high spatial resolution imaging (of MRI and CT techniques) [27]. This work reports the synthesis of β -NaGdF₄ nanocrystals doped with Yb³⁺ and Er³⁺ (NaGdF₄:Yb/Er). They were synthesized as core nanocrystals and were subsequently coated with a shell of NaGdF₄:Nd³⁺ (Core-shell nanocrystals).

EXPERIMENTAL

Chemicals: Gadolinium Oxide (99.9%), Ytterbium(III) acetate hydrate (99.99%), Erbium Oxide (99.9%), Neodymium(III) acetate hydrate (99.9%), Sodium hydroxide ($\geq 97\%$), 1-octadecene (90%), oleic acid (90%), Ammonium fluoride ($\geq 99.9\%$), Acetic Acid glacial ($\geq 99.85\%$) and cyclohexane were purchased from Sigma-Aldrich. All chemicals were used without any further purification.

Synthesis of Gadolinium acetate hydrate

Gadolinium acetate hydrate was prepared following the procedure reported by Dühnen and Haase for the synthesis of gadolinium trifluoroacetate (Gd-TFA) [28]. Gd₂O₃ (3.63 g, 10 mmol) was added to a solution of 80 mmol of concentrated acetic acid (4.6 mL) and 12.4 mL of distilled water. The solution was heated under reflux (at ca. 90 °C) for 60 min to dissolve the Gd₂O₃ completely. After removing excess acetic acid and water with a rotary evaporator at 50 °C, Gadolinium acetate hydrate was obtained as a white powder. Subsequently the powder was dried at 60 °C for 17 h.

Synthesis of Erbium acetate tetra hydrate

Erbium acetate tetra hydrate was prepared according to the procedure reported by Mondry and Bukietyńska [29]. Er_2O_3 was dissolved in concentrated acetic acid and alkalinized to pH = 4 using deionized water. The well-shaped crystals of Erbium acetate tetra hydrate were formed during very slow evaporation.

Synthesis of β -NaGdF₄ nanocrystals

β -NaGdF₄ nanocrystals were synthesized according to the procedure reported by Tian and his coworkers with a slight modification [30]. Gadolinium acetate hydrate (0.4 mmol) was added to a 50-mL three-neck round-bottom containing oleic acid (4 mL) and 1-octadecene (6 mL). The mixture was heated to 150 °C for 1 h under nitrogen atmosphere. It was then allowed to cool down to 50 °C. After that, a methanol solution (5.4 mL) of NH_4F (1.36 mmol) and NaOH (1 mmol) was added and stirred at 50 °C for 30 min under nitrogen atmosphere. To remove the methanol, the reaction mixture was heated at 100 °C for 30 min in vacuum. Subsequently, the temperature was raised to 290 °C and maintained for 1.5 h under N_2 atmosphere with continuous magnetic stirring. The solution was then allowed to cool to room temperature. The nanocrystals were precipitated by the addition of ethanol and isolated via centrifugation. The as-precipitated nanocrystals were washed several times with ethanol and dried in air at 70 °C overnight. The nanocrystals were collected and dispersed in 4 mL of cyclohexane for further characterization.

Synthesis of β -NaGdF₄:Yb/Er core nanocrystals

The lanthanide-doped NaGdF₄ core and core-shell nanocrystals were synthesized by a thermal decomposition method as described by Zheng *et al.* [31] with slight modification. In a typical procedure, 0.4 mmol of $\text{Gd}(\text{CH}_3\text{COO})_3 \cdot \text{H}_2\text{O}$, 0.392 mmol of $\text{Yb}(\text{CH}_3\text{COO})_3 \cdot 4\text{H}_2\text{O}$ and 0.008 mmol of $\text{Er}(\text{CH}_3\text{COO})_3 \cdot 4\text{H}_2\text{O}$ were mixed with 8 mL of oleic acid (OA) and 12 mL of 1-octadecene in a 50 mL three-neck round-bottom flask. The mixture was then heated to 150 °C and kept there for 60 min in the presence of continuous N_2 flow and magnetic stirring, resulting in a clear yellowish-brown solution. After cooling the solution to room temperature, 10 mL of methanol solution containing 2.72 mmol NH_4F and 2 mmol NaOH was added and the solution was stirred again at 50 °C for 30 min to remove methanol. Thereafter, the solution was heated to 290 °C and maintained for 90 min. After cooling down to room temperature, the nanocrystals were precipitated by addition of ethanol, and collected by centrifugation. They were washed with ethanol several times, and finally re-dispersed in cyclohexane.

Synthesis of β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd core-shell nanocrystals

For shell growth (β -NaGdF₄:Nd), 0.788 mmol of $\text{Gd}(\text{CH}_3\text{COO})_3 \cdot \text{H}_2\text{O}$ and 0.012 mmol of $\text{Nd}(\text{CH}_3\text{COO})_3 \cdot \text{H}_2\text{O}$ were added to a 50 mL three-neck round-bottom flask containing 8 mL of OA and 12 mL of 1-octadecene and heated to 150 °C under N_2 flow and magnetic stirring for 60 min; likewise a clear yellowish-brown solution was developed. After cooling down to 80 °C, 0.4 mmol β -NaGdF₄:Yb/Er nanocrystal seeds (core) in 6 mL of cyclohexane was added to the above solution. After complete evaporation of cyclohexane, 10 mL methanol solution containing 2.72 mmol NH_4F and 2 mmol NaOH was added. The mixture was then stirred at 50 °C for 30 min. In a similar fashion, after the evaporation of methanol the solution temperature was raised to 290 °C and maintained for 90 min under N_2 flow with vigorous stirring. The core-shell nanocrystals were obtained finally at room temperature with the same method as described above.

RESULTS AND DISCUSSION

All nanocrystals were characterized by Dynamic light scattering (DLS), X-ray Diffractometry (XRD) and Infrared (IR) analyses. DLS measurements were performed with a Zeta sizer Nano-Zs (Malvern, UK) at room temperature. The crystal structure of as-synthesized UCNPs was analyzed using D2 phase XRD (Bruker, Germany) equipped with PW3020 goniometer (Theta/2theta), 3kW generator (Cu anode), Nickel filter and operated using K-alpha1 ($\approx 1.54 \text{ \AA}$).

Dynamic Light scattering Study

The results of host (β -NaGdF₄), core (β -NaGdF₄:Yb/Er) and core-shell (β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd) nanocrystals confirmed their size to be 19.69 nm, 12.35 nm and 26.62 nm respectively (Fig. 1-3).

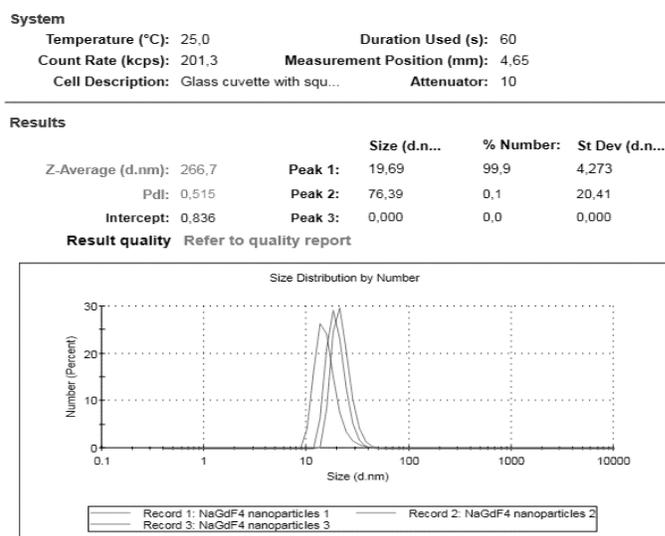


Figure 1. Dynamic Light Scattering Measurement (DLS) of Host (β -NaGdF₄), (Number vs Size).

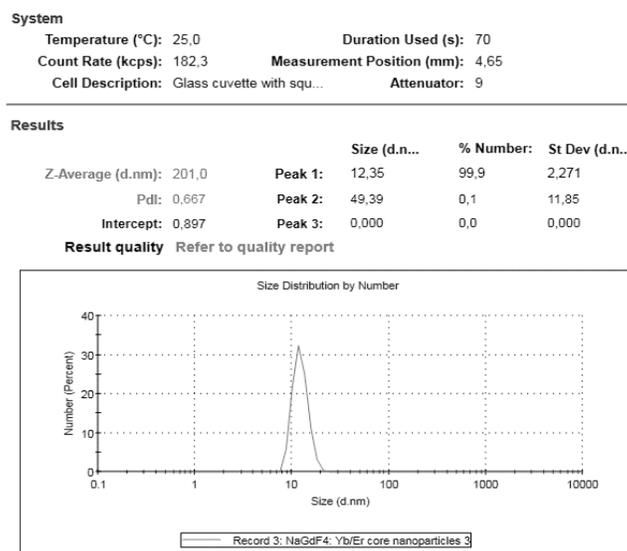


Figure 2. Dynamic Light Scattering Measurement (DLS) of core (NaGdF₄:Yb/Er), (Number vs Size).

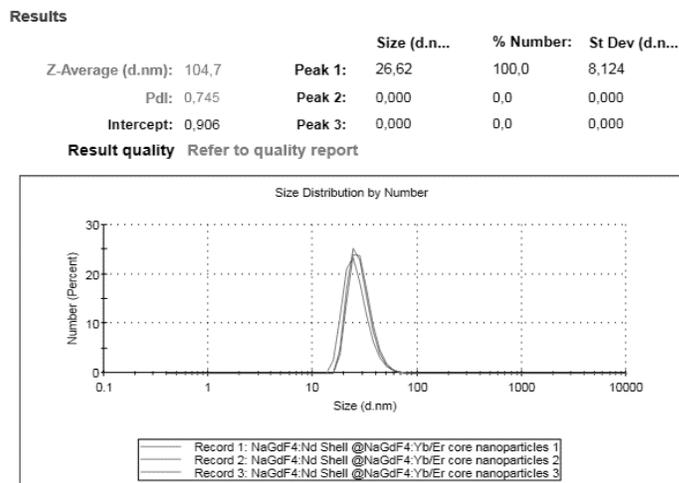


Figure 3. Dynamic Light Scattering Measurement (DLS) of β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd core-shell nanocrystals (Number vs Size).

XRD Study

The X-ray diffraction spectrum of the host β -NaGdF₄ nanocrystals is given in fig. 4. It shows well defined peaks. The X-ray diffraction reveals that the synthesized sample β -NaGdF₄ has some regular patterns which are due to the presence of β -NaGdF₄ crystals. The spectral lines present in the β -NaGdF₄ are summarized in Table. Values in Table are generated in Origin® 8.0, whereas thickness “t” and “d-spacing” of diffraction peaks are calculated using MS-Excel 2007.

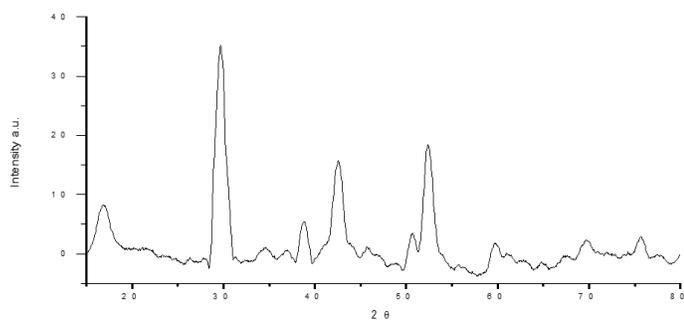


Figure 4. XRD pattern of β -NaGdF₄ nanocrystals

Table 1. XRD spectral lines data of β -NaGdF₄ nanocrystals

S.No	2 θ	Height From Base	Width at half height	d-spacing (Å)	Thickness “t” (nm)
1	17	4.07	1.67	5.27	4.98
2	30	15.66	1.24	3.00	6.87
3	43	7.81	1.22	2.12	7.22
4	52	8.71	1.11	1.74	8.23

Here in Table 1, the 2 θ , height of the peak from the base, width of the peak at half height, d-spacing (Calculated through Bragg’s equation) and their respective thickness (calculated through Scherer’s equation) are given. From Table one can easily observe that 2 θ values for major reflections ranges from 17° to 52°. Using the Bragg’s equation, the d-spacing in Å was calculated where it ranges from 1.74 Å to a maximum of 5.27 Å. The thickness of the crystal lattice was calculated by Scherer’s equation. The results indicate that crystal thickness ranges from 4.98 to 8.23 nm.

From figure 5, X-ray diffraction spectrum for the sample β -NaGdF₄:Yb/Er core may be observed. Well defined peaks indicate the presence of some repeated units in the sample which suggests that the sample β -NaGdF₄:Yb/Er is crystalline in nature. In Table 2, some well resolved reflections are summarized where 2 θ values ranges from 17° to 56° which through Bragg’s equation suggests that the spacing between the two reflections faces ranges from 1.64 Å to 5.22 Å. Applying Scherer’s equation, it was revealed that thickness of the crystals present in sample of β -NaGdF₄:Yb/Er core nanocrystals ranges from 0.94 nm to 8.66 nm.

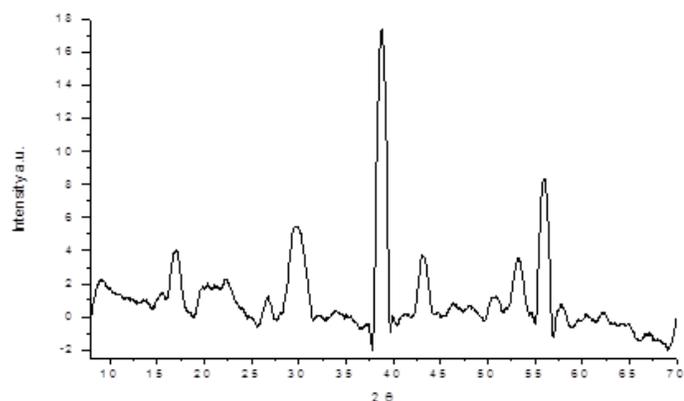


Figure 5. X-ray Diffraction (XRD) measurement of core (β -NaGdF₄:Yb/Er) nanocrystals

Table 2 XRD spectral lines data of β -NaGdF₄:Yb/Er nanocrystal.

S.No	2 θ	Height From Base	Width at half height	d-spacing (Å)	Thickness “t” (nm)
1	17	1.84	8.82	5.22	0.94
2	30	2.62	1.88	3.01	4.51
3	39	7.49	1.08	2.32	8.01
4	43	1.72	1.12	2.10	7.86
5	53	1.66	1.22	1.72	7.50
6	56	3.63	1.07	1.64	8.66

The X-ray diffraction spectrum for the sample of β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd core-shell is given in figure 6.

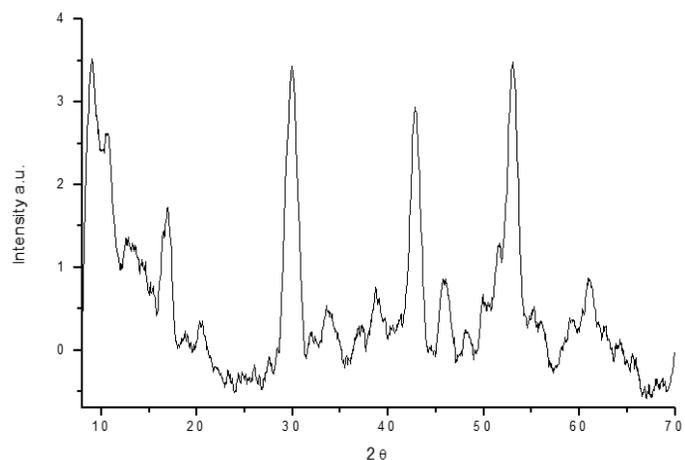


Figure 6. XRD pattern of β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd core-shell nanocrystals.

Well defined peaks indicate regular patterns in the core shell which are due to the presence of β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd crystals. The spectral lines are summarized in Table 3. Values in Table are generated in Origin® 8.0, whereas thickness “t” and “d-spacing” of diffraction peaks are calculated using MS-Excel 2007. In Table 3, the 2 θ values range from 17° to 56°. Applying Bragg’s equation the d-spacing between different faces is calculated which ranges from 1.54 Å to 9.81 Å.

Table 3. XRD spectral lines data of of β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd core-shell nanocrystals.

S.No	2 θ	Height From Base	Width at half height	d-spacing (Å)	Thickness “t” (nm)
1	09	1.10	1.93	9.81	4.26
2	17	0.83	1.27	5.21	6.53
3	30	1.56	1.42	2.98	5.98
4	39	0.33	1.08	2.31	8.06
5	43	1.42	1.23	2.10	7.17
6	46	0.02	0.56	1.97	15.92
7	52	0.11	1.92	1.76	4.76
8	53	1.72	1.36	1.72	6.74
9	60	0.38	1.44	1.54	6.58

These experimental values further reveal that the thickness of these crystalline structures is from 4.26 nm to 15.92 nm.

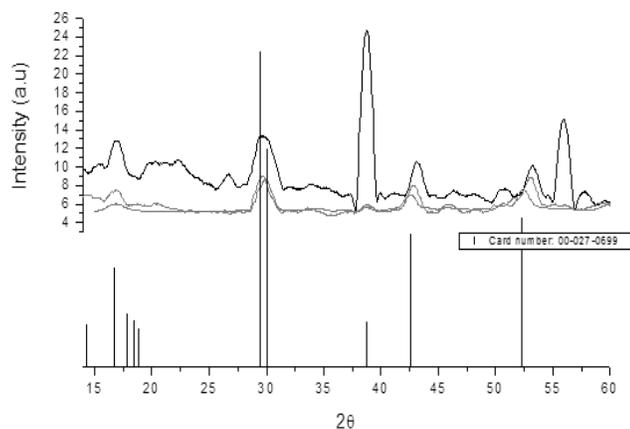


Figure 7. XRD spectra of standard reference pattern of β -NaGdF₄ (red) (JCPDS 27-0699), core β -NaGdF₄:Yb/Er (black) and core-shell β -NaGdF₄:Yb/Er@ β -NaGdF₄:Nd (blue) nanocrystals

From figure 7, it may be observed that the diffractogram of all crystals (core and core shell crystals) are alike. The positions of well-defined diffraction peaks are consistent with that of reference crystals of NaGdF₄ in hexagonal structure (JCPDS 27-0699).

FT-IR Studies

Fourier transform infrared (FTIR) spectra of the prepared NaGdF₄ (host), NaGdF₄:Yb/Er (Core) and NaGdF₄:Yb/Er@NaGdF₄:Nd (Core@shell) nanocrystals are shown in figure 8, where peaks at 2927 and 2852 cm⁻¹ indicate the characteristic stretching of -CH₂ group of the oleic acid attached to the NaGdF₄ nanocrystals. Peaks at 1565 cm⁻¹ and 1464 cm⁻¹ are attributed to -COOH stretching, whereas peaks at 3416 cm⁻¹ and 1551 cm⁻¹ are assigned to -OH and C=O stretching respectively. Furthermore, two peaks appeared at 1402 cm⁻¹ and 1093 cm⁻¹ indicate that -OH, C-OH, and -COOH groups are present at the surface of the synthesized nanocrystals.

The FTIR spectrum of the nanocrystals was similar to the spectra given in the literature [32, 33] where the presence of -OH and -COOH indicate that some hydrophilic character may be one of the prominent properties of the synthesized nanocrystals making them favorable for applications in the field of biomedicines and alike.

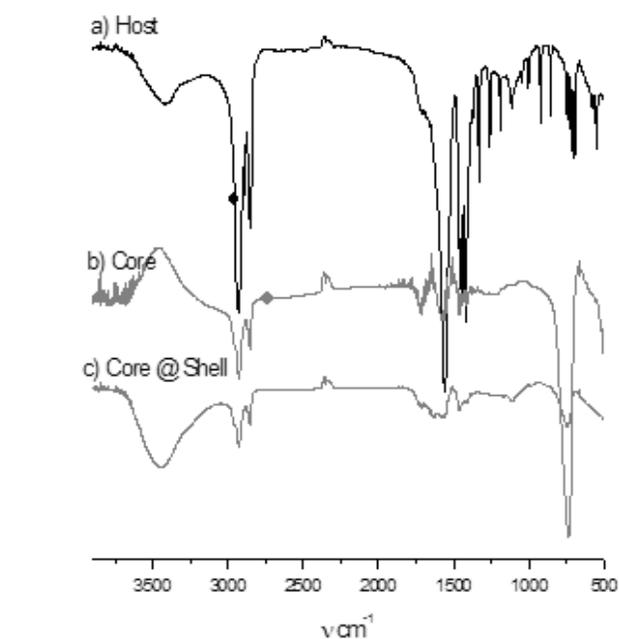
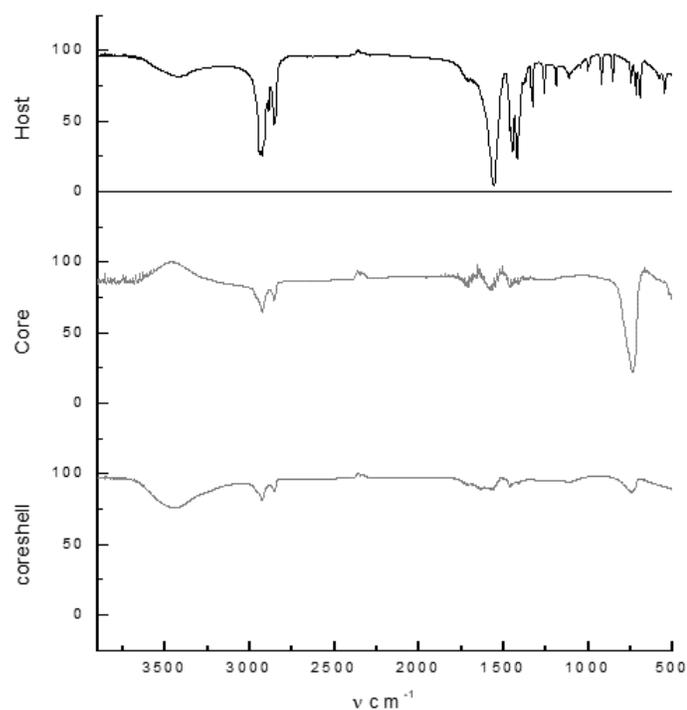


Figure 8. (a) Host: NaGdF₄ b) Core: NaGdF₄:Yb/Er c) Core@shell: NaGdF₄:Yb/Er@NaGdF₄:Nd

CONCLUSIONS

This work reports the synthesis of β -NaGdF₄ nanocrystals. These nanocrystals were doped with Yb and Er to produce core nanocrystals (β -NaGdF₄:Yb/Er). A layer of β -NaGdF₄:Nd was coated as shell over the core nanocrystals to produce core-shell nanocrystals as β -NaGdF₄:Yb/Er@ β -NaGdF₄:Yb/Er. β -NaGdF₄ nanomaterials doped with Yb and Er (green and red emission) are well known up-converting materials. Therefore, core-shell nanocrystals prepared in this work can be investigated as up-converting materials. They are expected to be useful for X-ray tomography because gadolinium is a heavy element, capable of strongly absorbing X-rays. Gadolinium-based nanomaterials are also applied as MRI contrast agents because Gd increases water protons relaxation rate. Thus, these nanocrystals are likely to be suitable contrast agents for photoluminescence, MRI and X-ray tomography. However, further research is needed for their *in vivo* investigations.

CONFLICTS OF INTEREST

The authors declare no conflicts of interest.

ACKNOWLEDGMENTS

The authors are thankful to Brazilian Foundation Agency (CNPq) for providing financial support to execute this project (400662/2013-2).

REFERENCES

1. A. K. Buck, S. Nekolla, S. Ziegler, A. Beer, B. J. Krause, K. Herrmann, K. Scheidhauer, H.-J. Wester, E. J. Rummeny M. J. J. N. M. Schwaiger, J. Nucl. Med. 49, 1305 (2008)
2. A. Galperin, D. Margel, J. Baniel, G. Dank, H. Biton S. Margel, Biomaterials 28, 4461 (2007)
3. O. M. Merkel, D. Librizzi, A. Pfestroff, T. Schurrat, M. Béhé T. Kissel, Bioconj. Chem. 20, 174 (2008)
4. B. J. Pichler, H. F. Wehr M. S. J. J. o. N. M. Judenhofer, J. Nucl. Mater. 49, 5S (2008)
5. B. J. Pichler, A. Kolb, T. Nägele H.-P. Schlemmer, J. Nucl. Med. 51, 333 (2010)
6. R. Bar-Shalom, N. Yefremov, L. Guralnik, D. Gaitini, A. Frenkel, A. Kuten, H. Altman, Z. Keidar O. Israel, J. Nucl. Med. 44, 1200 (2003)
7. A. Dirksen, S. Langereis, B. F. de Waal, M. H. van Genderen, T. M. Hackeng E. Meijer, Chem. Commun. 2811 (2005).

8. J.-L. Bridot, A.-C. Faure, S. Laurent, C. Riviere, C. Billotey, B. Hiba, M. Janier, V. Jossierand, J.-L. Coll L. Vander Elst, J. Am. Chem. Soc. 129, 5076 (2007)
9. J. Liu, W. Bu, S. Zhang, F. Chen, H. Xing, L. Pan, L. Zhou, W. Peng J. Shi, Chem. A Eur. J. 18, 2335 (2012)
10. S. Adams, R. P. Baum, T. Stuckensen, K. Bitter G. Hör, Eur. J. Nucl. Med. 25, 1255 (1998)
11. H. U. Rashid, K. Yu J. Zhou, J. Struct. Chem. 54, 223 (2013)
12. R. H. U., M. N. Umar, K. Khan, M. N. Anjum M. Yaseen, J. Struct. Chem. 55, 910 (2014)
13. H. U. Rashid, K. Khan, M. Yaseen, W. Hassan M. N. UMAR, Rom. J. Chem. 59, 27 (2014)
14. F. Zhang, in *Photon Upconversion Nanomaterials*, Springer, 2015, pp. 187.
15. J. Cheon J.-H. Lee, Acc. Chem. Res. 41, 1630 (2008)
16. W. J. Mulder, A. W. Griffioen, G. J. Strijkers, D. P. Cormode, K. Nicolay Z. A. Fayad, Nanomedicine 2, 307 (2007)
17. C. Bremer R. Weissleder, Acad. Radiol. 8, 15 (2001)
18. T. Persigehl, W. Heindel C. Bremer, Abdom. Imaging 30, 343 (2005)
19. D. Sosnovik R. Weissleder, in *Imaging in Drug Discovery and Early Clinical Trials*, Springer, 2005, pp. 83.
20. F. Liu, X. He, L. Liu, H. You, H. Zhang Z. Wang, Biomaterials 34, 5218 (2013)
21. D. K. Chatterjee, M. K. Gnanasammandhan Y. Zhang, Small 6, 2781 (2010)
22. J. Zhou, X. Zhu, M. Chen, Y. Sun F. Li, Biomaterials 33, 6201 (2012)
23. C. Wang, H. Tao, L. Cheng Z. Liu, Biomaterials 32, 6145 (2011)
24. K. W. Krämer, D. Biner, G. Frei, H. U. Güdel, M. P. Hehlen S. R. Lüthi, Chem. Mater. 16, 1244 (2004)
25. M. He, P. Huang, C. Zhang, H. Hu, C. Bao, G. Gao, R. He D. Cui, Adv. Funct. Mater. 21, 4470 (2011)
26. S. Zeng, M.-K. Tsang, C.-F. Chan, K.-L. Wong J. Hao, Biomaterials 33, 9232 (2012)
27. Y. F. Wang, L. D. Sun, J. W. Xiao, W. Feng, J. C. Zhou, J. Shen C. H. Yan, Chem. A Eur. J. 18, 5558 (2012)
28. S. Dühnen M. Haase, Chem. Mater. 27, 8375 (2015)
29. A. Mondry K. J. A. P. P.-S. A. G. P. Bukietynska, Acta Physica Polonica-Series A General Physics 90, 233 (1996)
30. J. Tian, X. Zeng, X. Xie, S. Han, O.-W. Liew, Y.-T. Chen, L. Wang X. Liu, J. Am. Chem. Soc. 137, 6550 (2015)
31. S. Zheng, W. Chen, D. Tan, J. Zhou, Q. Guo, W. Jiang, C. Xu, X. Liu J. Qiu, Nanoscale 6, 5675 (2014)
32. X. Sun Y. Li, Angew. Chem. 116, 607 (2004)
33. X. Liu, L. Wang, Z. Wang Z. Li, J. Mater. Res. 26, 82 (2011)